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(57) Abstract <p>Phosphopeptides containing the Ser(P) cluster sequence motif Ser(P)-Ser(P)-Ser(P)-Glu-Glu- can stabilise their own weight in amorphous calcium phosphate (ACP) [Ca<sub>3</sub>(PO<sub>4</sub>)<sub>1.87</sub>(HPO<sub>4</sub>)<sub>0.2</sub>xH<sub>2</sub>O] and amorphous calcium fluoride phosphate (ACFP) [Ca<sub>3</sub>(PO<sub>4</sub>)<sub>3</sub> F x H<sub>2</sub>O]. The amorphous phases stabilised by the phosphopeptides are an excellent delivery vehicle to co-localise Ca, F, and phosphate at the tooth surface in a slow-release amorphous form producing superior anticaries efficacy. These amorphous phases stabilised by the phosphopeptides also have utility as dietary supplements to increase calcium bioavailability and to help prevent diseases associated with calcium deficiencies.</p>			

## CALCIUM PHOSHOPEPTIDE COMPLEXES

The present invention relates to novel complexes in which amorphous calcium phosphates are stabilised by phosphopeptides. These complexes have anti-cariogenic effects, and may also be used as dietary supplements to increase calcium bioavailability and to heal or prevent diseases associated with calcium deficiencies. Methods of making the complexes of the invention and of treatment or prevention of dental caries, calcium malabsorption, and bone diseases are also provided.

### BACKGROUND

#### 10 Dental Caries

Dental caries is initiated by the demineralisation of hard tissue on the teeth by organic acids produced from fermentation of dietary sugar by dental plaque odontopathogenic bacteria.

Even though the prevalence of dental caries has decreased through the 'use of fluoride in most developed countries, the disease remains a major public health problem. The estimated economic burden of treating dental caries in Australia in 1991 was \$471 million, being higher than that for other diet-related diseases including coronary heart disease, hypertension or stroke.

In developing countries where the availability of industrialised food products is increasing, prevalence of dental caries is also increasing. Recent studies have highlighted a number of socio-demographic variables associated with the risk of developing caries; high risk is associated with ethnicity and low socio-economic status. The level of high-risk individuals has remained constant even though the overall severity and prevalence of disease in the community has decreased. Dental caries is therefore, still a major public health problem, particularly in ethnic and lower socio-economic groups. This highlights the need for a non-toxic, anticariogenic agent that could supplement the effects of fluoride to further lower the incidence of dental caries. An agent which would reduce the dose of fluoride required to reduce the incidence of caries would be particularly desirable in view of community anxiety about fluoride, and in view

of the fact that fluorosis can develop even at currently used doses.

The food group most recognised as exhibiting anticaries activity is dairy products (milk, milk concentrates, powders and cheeses). US Patent No. 5,130,123 discloses the component responsible for this anticariogenic activity as casein. However, the use of casein as an anticariogenic agent is precluded by adverse organoleptic properties and the very high levels required for activity.

Preliminary investigations determined that tryptic casein phosphopeptides contributed to the anticariogenic activity and this was made subject of US Patent No. 5,015,628. In particular, peptides Bos  $\alpha_1$ -casein X-5P (f59-79) [1], Bos  $\beta$ -casein X-4P (f1-25) [2], Bos  $\alpha_2$ -casein X-4P (f46-70) [3] and Bos  $\alpha_2$ -casein X-4P (f1-21) [4] were disclosed in US Patent 5,015,628 as follows:

[1] Gln<sup>59</sup>-Met-Glu-Ala-Glu-Ser(P)-Ile-Ser(P)-Ser(P)-Ser(P)-Glu-Glu-Ile-Val-Pro-Asn-Ser(P)-Val-Glu-Gln-Lys<sup>79</sup>.  $\alpha_1(59-79)$

[2] Arg<sup>1</sup>-Glu-Leu-Glu-Glu-Leu-Asn-Val-Pro-Gly-Glu-Ile-Val-Glu-Ser(P)-Leu-Ser(P)-Ser(P)-Ser(P)-Glu-Glu-Ser-Ile-Thr-Arg<sup>25</sup>.  $\beta(1-25)$

[3] Asn<sup>46</sup>-Ala-Asn-Glu-Glu-Glu-Tyr-Ser-Ile-Gly-Ser(P)-Ser(P)-Ser(P)-Glu-Glu-Ser(P)-Ala-Glu-Val-Ala-Thr-Glu-Glu-Val-Lys<sup>70</sup>.  $\alpha_2(46-70)$

[4] Lys<sup>1</sup>-Asn-Thr-Met-Glu-His-Val-Ser(P)-Ser(P)-Ser(P)-Glu-Glu-Ser-Ile-Ile-Ser(P)-Gln-Glu-Thr-Tyr-Lys<sup>21</sup>.  $\alpha_2(1-21)$

The preliminary determination of the above phosphopeptides for use in combination with CaHPO<sub>4</sub> and hydroxyapatite provided novel peptides having anticariogenic properties. However, subsequent investigations have determined that the Ser(P) cluster sequence motif within the previous disclosed phosphopeptides have the unexpected ability to stabilize their own weight in amorphous calcium phosphate. The ability of the above phosphopeptides and

in particular the Ser(P) motif to stabilise amorphous calcium phosphate was quite unexpected and neither disclosed or taught in any publications known to the Applicants. We have now found that the amorphous form of calcium phosphate  $\text{Ca}_3(\text{PO}_4)_{1.87}(\text{HPO}_4)_{0.2}x\text{H}_2\text{O}$  where  $x \geq 1$  stabilised by the casein phosphopeptides is the most soluble, basic form of non-crystalline calcium phosphate and a superior form of calcium phosphate which prevents caries and increases calcium bioavailability. Amorphous calcium phosphate (ACP) must be formed by careful titration of Ca ions (eg  $\text{CaCl}_2$ ) and phosphate ions (eg  $\text{NaHPO}_4$ ) while maintaining the pH above 7 (preferably 9.0) in the presence of the phosphopeptide. As the ACP is formed, the phosphopeptide binds to the nascent nuclei and stabilises the ACP as a phosphopeptide-ACP complex. Without the phosphopeptide, the ACP will precipitate out of solution and transform within minutes into the most stable calcium phosphate phase, crystalline hydroxyapatite (HA). HA, by being insoluble has limited anticariogenic activity and presents calcium in a poorly bioavailable form. The acidic phase of calcium phosphate  $\text{CaHPO}_4$ , while certainly being more soluble than hydroxyapatite, is poorly bound by the phosphopeptide and poorly localised at the tooth surface and therefore also has limited anticariogenic activity. The unexpected ability of the aforementioned phosphopeptides and in particular Ser(P) cluster motif to stabilize amorphous calcium phosphate was not disclosed or taught in US Patent No. 5,015,628 and provides for the first time a reliable and effective method of producing a stabilized amorphous calcium phosphate complex having distinct and novel advantages in calcium treatments and delivery. US Patent No. 5,015,628 does not disclose the unique amorphous calcium fluoride phosphate phase  $\text{Ca}_3(\text{PO}_4)_2\text{F} \cdot x\text{H}_2\text{O}$  where  $x \geq 1$  which we have now found to be stabilised by the above phosphopeptides and can be localised at the tooth surface to provide superior anticaries efficacy. This unexpected ability to stabilize amorphous calcium phosphate forms the basis of the instant invention.

### SUMMARY OF THE INVENTION

In one aspect, the invention provides a stable calcium phosphate complex, comprising amorphous calcium phosphate or a derivative thereof stabilised by a phosphopeptide, wherein said phosphopeptide comprises the sequence -  
5 Ser(P)-Ser(P)-Ser(P)-Glu-Glu-.

The amorphous calcium phosphate (ACP) is preferably of the formula  $[\text{Ca}_3(\text{PO}_4)_{1.87} (\text{HPO}_4)_{0.2} \times \text{H}_2\text{O}]$  wherein  $x \geq 1$ . The calcium phosphate derivative may be calcium fluoride phosphate having the formula  $[\text{Ca}_8(\text{PO}_4)_5 \text{F} \times \text{H}_2\text{O}]$  wherein  $x \geq 1$ , providing amorphous calcium fluoride phosphate (ACFP).

10 The phosphopeptide (PP) may be from any source; it may be obtained by tryptic digestion of casein or other phospho-acid rich proteins such as phosphitin, or by chemical or recombinant synthesis, provided that it comprises the core sequence -Ser(P)-Ser(P)-Ser(P)-Glu-Glu-. The sequence flanking this core sequence may be any sequence. However, those flanking sequences in  
15  $\alpha_1(59-79)$  [1],  $\beta(1-25)$  [2],  $\alpha_2(46-70)$  [3] and  $\alpha_2(1-21)$  [4] are preferred. The flanking sequences may optionally be modified by deletion, addition or conservative substitution of one or more residues. The amino acid composition and sequence of the flanking region are not critical as long as the conformation of the peptide is maintained and that all phosphoryl and carboxyl groups  
20 interacting with calcium ions are maintained as the preferred flanking regions appear to contribute to the structural action of the motif.

The complex formed preferably has the formula  $[(\text{PP})(\text{CP})_8]_n$  where n is equal to or greater than 1, for example, 6. The complex formed may be a colloidal complex.

25 The phosphopeptide binds to the ACP cluster to produce a metastable solution in which growth of ACP to a size that initiates nucleation and precipitation is prevented. In this way, calcium and other ions such as fluoride ions can be localised, for instance at a surface on a tooth to prevent demineralisation and prevent formation of dental caries.

Thus, in a second aspect, the invention provides a stable calcium phosphate complex as described above, which complex acts as a delivery vehicle that co-localises ions including, but not limited to calcium, fluoride and phosphate ions at a target site. In a preferred embodiment, the complex is in a slow-release amorphous form that produces superior anti-caries efficacy.

In a particularly preferred embodiment of the invention, the stable calcium complex is incorporated into dentifrices such as toothpaste, mouth washes or formulations for the mouth to aid in the prevention and/or treatment of dental caries or tooth decay. The calcium complex may comprise 0.05-50% by weight of the composition, preferably 1.0-50%. For oral compositions, it is preferred that the amount of the CPP-ACP and/or CPP-ACFP administered is 0.05 - 50% by weight, preferably 1.0% - 50% by weight of the composition. The oral composition of this invention which contains the above-mentioned agents may be prepared and used in various forms applicable to the mouth such as dentifrice including toothpastes, toothpowders and liquid dentifrices, mouthwashes, troches, chewing gums, dental pastes, gingival massage creams, gargle tablets, dairy products and other foodstuffs. The oral composition according to this invention may further include additional well known ingredients depending on the type and form of a particular oral composition.

In certain highly preferred forms of the invention the oral composition may be substantially liquid in character, such as a mouthwash or rinse. In such a preparation the vehicle is typically a water-alcohol mixture desirably including a humectant as described below. Generally, the weight ratio of water to alcohol is in the range of from about 1:1 to about 20:1. The total amount of water-alcohol mixture in this type of preparation is typically in the range of from about 70 to about 99.9% by weight of the preparation. The alcohol is typically ethanol or isopropanol. Ethanol is preferred.

The pH of such liquid and other preparations of the invention is generally in the range of from about 5 to about 9 and typically from about 7.0 - 9.0. The pH can be controlled with acid (e.g. citric acid or benzoic acid) or base (e.g.

sodium hydroxide) or buffered (as with sodium citrate, benzoate, carbonate, or bicarbonate, disodium hydrogen phosphate, sodium dihydrogen phosphate, etc).

In other desirable forms of this invention, the oral composition may be substantially solid or pasty in character, such as toothpowder, a dental tablet or  
5 a toothpaste (dental cream) or gel dentifrice. The vehicle of such solid or pasty oral preparations generally contains dentally acceptable polishing material. Examples of polishing materials are water-insoluble sodium metaphosphate, potassium metaphosphate, tricalcium phosphate, dihydrated calcium phosphate, anhydrous dicalcium phosphate, calcium pyrophosphate, magnesium  
10 orthophosphate, trimagnesium phosphate, calcium carbonate, hydrated alumina, calcined alumina, aluminum silicate, zirconium silicate, silica, bentonite, and mixtures thereof. Other suitable polishing material include the particulate thermosetting resins such as melamine-, phenolic, and urea-formaldehydes, and cross-linked polyepoxides and polyesters. Preferred polishing materials include  
15 crystalline silica having particle sized of up to about 5 microns, a mean particle size of up to about 1.1 microns, and a surface area of up to about 50,000 cm<sup>2</sup>/gm., silica gel or colloidal silica, and complex amorphous alkali metal aluminosilicate.

When visually clear gels are employed, a polishing agent of colloidal  
20 silica, such as those sold under the trademark SYLOID as Syloid 72 and Syloid 74 or under the trademark SANTOCEL as Santocel 100, alkali metal aluminosilicate complexes are particularly useful since they have refractive indices close to the refractive indices of gelling agent-liquid (including water and/or humectant) systems commonly used in dentifrices.

25 Many of the so-called "water insoluble" polishing materials are anionic in character and also include small amounts of soluble material. Thus, insoluble sodium metaphosphate may be formed in any suitable manner as illustrated by Thorpe's Dictionary of Applied Chemistry, Volume 9, 4th Edition, pp. 510-511. The forms of insoluble sodium metaphosphate known as Madrell's salt and  
30 Kurrol's salt are further examples of suitable materials. These metaphosphate

salts exhibit only a minute solubility in water, and therefore are commonly referred to as insoluble metaphosphates (IMP). There is present therein a minor amount of soluble phosphate material as impurities, usually a few percent such as up to 4% by weight. The amount of soluble phosphate material, which is  
5 believed to include a soluble sodium trimetaphosphate in the case of insoluble metaphosphate, may be reduced or eliminated by washing with water if desired. The insoluble alkali metal metaphosphate is typically employed in powder form of a particle size such that no more than 1% of the material is larger than 37  
microns.

10 The polishing material is generally present in the solid or pasty compositions in weight concentrations of about 10% to about 99%. Preferably, it is present in amounts from about 10% to about 75% in toothpaste, and from about 70% to about 99% in toothpowder. In toothpastes, when the polishing material is silicious in nature, it is generally present in amount of about 10-30%  
15 by weight. Other polishing materials are typically present in amount of about 30-75% by weight.

In a toothpaste, the liquid vehicle may comprise water and humectant typically in an amount ranging from about 10% to about 80% by weight of the preparation. Glycerine, propylene glycol, sorbitol and polypropylene glycol  
20 exemplify suitable humectants/carriers. Also advantageous are liquid mixtures of water, glycerine and sorbitol. In clear gels where the refractive index is an important consideration, about 2.5 - 30% w/w of water, 0 to about 70% w/w of glycerine and about 20-80% w/w of sorbitol are preferably employed.

Toothpaste, creams and gels typically contain a natural or synthetic  
25 thickener or gelling agent in proportions of about 0.1 to about 10, preferably about 0.5 to about 5% w/w. A suitable thickener is synthetic hectorite, a synthetic colloidal magnesium alkali metal silicate complex clay available for example as Laponite (e.g. CP, SP 2002, D) marketed by Laporte Industries Limited. Laponite D is, approximately by weight 58.00% SiO<sub>2</sub>, 25.40% MgO,  
30 3.05% Na<sub>2</sub>O, 0.98% Li<sub>2</sub>O, and some water and trace metals. Its true specific



gravity is 2.53 and it has an apparent bulk density of 1.0 g/ml at 8% moisture.

Other suitable thickeners include Irish moss, iota carrageenan, gum tragacanth, starch, polyvinylpyrrolidone, hydroxyethylpropylcellulose, hydroxybutyl methyl cellulose, hydroxypropyl methyl cellulose, hydroxyethyl  
5 cellulose (e.g. available as Natrosol), sodium carboxymethyl cellulose, and colloidal silica such as finely ground Syloid (e.g. 244). Solubilizing agents may also be included such as humectant polyols such propylene glycol, dipropylene glycol and hexylene glycol, cellosolves such as methyl cellosolve and ethyl cellosolve, vegetable oils and waxes containing at least about 12 carbons in a  
10 straight chain such as olive oil, castor oil and petrolatum and esters such as amyl acetate, ethyl acetate and benzyl benzoate.

It will be understood that, as is conventional, the oral preparations are to be sold or otherwise distributed in suitable labelled packages. Thus, a jar of mouthrinse will have a label describing it, in substance, as a mouthrinse or  
15 mouthwash and having directions for its use; and a toothpaste, cream or gel will usually be in a collapsible tube, typically aluminium, lined lead or plastic, or other squeeze, pump or pressurized dispenser for metering out the contents, having a label describing it, in substance, as a toothpaste, gel or dental cream.

Organic surface-active agents are used in the compositions of the present  
20 invention to achieve increased prophylactic action, assist in achieving thorough and complete dispersion of the active agent throughout the oral cavity, and render the instant compositions more cosmetically acceptable. The organic surface-active material is preferably anionic, nonionic or ampholytic in nature and preferably does not interact with the active agent. It is preferred to employ  
25 as the surface-active agent a detergent material which imparts to the composition detergent and foaming properties. Suitable examples of anionic surfactants are water-soluble salts of higher fatty acid monoglyceride monosulfates, such as the sodium salt of the monosulfated monoglyceride of hydrogenated coconut oil fatty acids, higher alkyl sulfates such as sodium lauryl sulfate, alkyl aryl  
30 sulfonates such as sodium dodecyl benzene sulfonate, higher alkylsulfo-acetates,

higher fatty acid esters of 1,2-dihydroxy propane sulfonate, and the substantially saturated higher aliphatic acyl amides of lower aliphatic amino carboxylic acid compounds, such as those having 12 to 16 carbons in the fatty acid, alkyl or acyl radicals, and the like. Examples of the last mentioned amides are N-lauroyl  
5 sarcosine, and the sodium, potassium, and ethanolamine salts of N-lauroyl, N-myristoyl, or N-palmitoyl sarcosine which should be substantially free from soap or similar higher fatty acid material. The use of these sarconite compounds in the oral compositions of the present invention is particularly advantageous since these materials exhibit a prolonged marked effect in the  
10 inhibition of acid formation in the oral cavity due to carbohydrates breakdown in addition to exerting some reduction in the solubility of tooth enamel in acid solutions. Examples of water-soluble nonionic surfactants suitable for use are condensation products of ethylene oxide with various reactive hydrogen-containing compounds reactive therewith having long hydrophobic chains (e.g.  
15 aliphatic chains of about 12 to 20 carbon atoms), which condensation products ("ethoxamers") contain hydrophilic polyoxyethylene moieties, such as condensation products of poly (ethylene oxide) with fatty acids, fatty alcohols, fatty amides, polyhydric alcohols (e.g. sorbitan monostearate) and polypropyleneoxide (e.g. Pluronic materials).

20 The surface active agent is typically present in amount of about 0.1-5% by weight. It is noteworthy, that the surface active agent may assist in the dissolving of the active agent of the invention and thereby diminish the amount of solubilizing humectant needed.

25 Various other materials may be incorporated in the oral preparations of this invention such as whitening agents, preservatives, silicones, chlorophyll compounds and/or ammoniated material such as urea, diammonium phosphate, and mixtures thereof. These adjuvants, where present, are incorporated in the preparations in amounts which do not substantially adversely affect the properties and characteristics desired.

Any suitable flavouring or sweetening material may also be employed. Examples of suitable flavouring constituents are flavouring oils, e.g. oil of spearmint, peppermint, wintergreen, sassafras, clove, sage, eucalyptus, marjoram, cinnamon, lemon, and orange, and methyl salicylate. Suitable  
5 sweetening agents include sucrose, lactose, maltose, sorbitol, xylitol, sodium cyclamate, perillartine, AMP (aspartyl phenyl alanine, methyl ester), saccharine, and the like. Suitably, flavour and sweetening agents may each or together comprise from about 0.1% to 5% more of the preparation.

In the preferred practice of this invention an oral composition according  
10 to this invention such as mouthwash or dentifrice containing the composition of the present invention is preferably applied regularly to the gums and teeth, such as every day or every second or third day or preferably from 1 to 3 times daily, at a pH of about 4.5 to about 9, generally about 7.0 to about 9, for at least 2 weeks up to 8 weeks or more up to a lifetime.

15 The compositions of this invention can also be incorporated in lozenges, or in chewing gum or other products, e.g. by stirring into a warm gum base or coating the outer surface of a gum base, illustrative of which may be mentioned jelutong, rubber latex, vinylite resins, etc., desirably with conventional plasticizers or softeners, sugar or other sweeteners or such as glucose, sorbitol  
20 and the like.

In another embodiment, the complex of the invention is formulated to form a dietary supplement preferably comprising 0.1-100% w/w, more preferably 1-50% w/w, most preferably 1-10% and particularly 2% w/w. The complex may also be incorporated into food products.

25 Accordingly, in a third aspect, the invention provides compositions including pharmaceutical compositions comprising the calcium complex as described together with a pharmaceutically-acceptable carrier. Such compositions may be selected from the group consisting of dental, anti-cariogenic compositions, therapeutic compositions and dietary supplements.  
30 Dental compositions or therapeutic compositions may be in the form of a gel,

liquid, solid, powder, cream or lozenge. Therapeutic compositions may also be in the form of tablets or capsules.

In a fourth aspect, there is provided a method of treating or preventing dental caries or tooth decay comprising the step of administering a complex or composition of the invention to the teeth or gums of a subject in need of such treatments. Topical administration of the complex is preferred.

In a fifth aspect, the invention relates to methods of treating one or more conditions related to calcium loss from the body, especially from the bones, calcium deficiency, calcium malabsorption, or the like. Examples of such conditions include, but are not limited to, osteoporosis and osteomalacia. In general any condition which can be improved by calcium bioavailability is contemplated.

In a sixth aspect, the invention also provides a method of producing a stable complex of calcium phosphate as described above, comprising the step of:

- (i) obtaining a solution of phosphopeptide having a pH of about 9.0;
- (ii) admixing (i) with solutions comprising calcium, and inorganic phosphate and optionally fluoride at a pH of about 9.0;
- (iii) filtering the mixture resulting from step (ii), and
- (iv) drying to obtain the said complex.

The complexes of the invention are useful as calcium supplements in subjects in need of stimulation of bone growth, for example subjects undergoing fracture repair, joint replacement, bone grafts, or craniofacial surgery.

These complexes are also useful as dietary supplements in subjects who for any reason, such as dietary intolerance, allergy, or religious or cultural factors, are unable or unwilling to consume dairy products in an amount sufficient to supply their dietary calcium requirements.

It will be clearly understood that, although this specification refers specifically to applications in humans, the invention is also useful for veterinary purposes. Thus in all aspects the invention is useful for domestic animals such

as cattle, sheep, horses and poultry; for companion animals such as cats and dogs; and for zoo animals.

### **DETAILED DESCRIPTION OF THE INVENTION**

The invention will now be described in detail by way of reference only  
5 to the following non-limiting Examples.

#### **Example 1 Preparation of CPP-ACP and CPP-ACFP**

##### ***A. Preparation of CPP-ACP***

A 10% w/v casein (Murray Goulburn, Victoria, Australia) or caseinate  
solution was prepared at pH 8.0 and then digested with trypsin at 0.2% w/w of  
10 the casein for 2h at 50°C with the pH controlled to  $8.0 \pm 0.1$  by NaOH  
addition. After digestion the solution was adjusted to pH 4.6 by the addition  
of HCl and the precipitate removed by centrifugation or microfiltration.  
However, the solution can also be clarified by microfiltration at pH 8.0 without  
acidification. The supernatant or microfiltrate was then adjusted to pH 9.0 with  
15 NaOH, then  $\text{CaCl}_2$  (1.6 M) and  $\text{Na}_2\text{HPO}_4$  (1 M) at pH 9.0 were added slowly  
( $\leq 1\%$  vol per min) with constant agitation with the pH held constant at  $9.0 \pm 0.1$   
by NaOH addition.  $\text{CaCl}_2$  and sodium phosphate were added to the final  
concentrations of 100 mM and 60 mM respectively. Following the addition of  
the calcium and phosphate solutions, the solution was microfiltered through a  
20 0.1 or 0.2  $\mu\text{m}$  microfilter (ceramic or organic) to concentrate the solution five  
fold. The retentate was then diafiltered with one to five volumes of distilled  
water. The retentate after diafiltration was spraydried to produce a white  
powder that was 50% CPP and 40% ACP and residue water. Analysis of the  
CPP of the CPP-ACP complex by reversed-phase HPLC, sequence analysis and  
25 mass spectrometry revealed that the only peptides that are capable of stabilising  
the amorphous calcium phosphate and retained during the microfiltration and  
diafiltration are Bos  $\alpha_{s1}$ -casein X-5P (f59-79) [1], Bos  $\beta$ -casein X-4P (f1-25)  
[2], Bos  $\alpha_{s2}$ -casein X-4P (f46-70) [3] and Bos  $\alpha_{s2}$ -casein X-4P (f1-21) [4] and  
truncated and heat modified forms of these peptides.

### *B. Preparation of CPP-ACFP*

A 10% w/v casein or caseinate solution was prepared at pH  $8.0 \pm 0.1$  and then digested with trypsin at 0.2% w/w of the casein for 2h at 50°C. After digestion the solution was adjusted to pH 4.6 by the addition of HCl and the precipitate removed by centrifugation or microfiltration. However the solution can also be clarified by microfiltration at pH 8.0 without acidification. The supernatant or microfiltrate was then adjusted to pH 9.0 with NaOH, then  $\text{CaCl}_2$  (1.6 M),  $\text{Na}_2\text{HPO}_4$  (1 M) at pH 9.0 and 200 mM NaF were added slowly ( $\leq 1\%$  vol per min) with constant agitation with the pH held constant at  $9.0 \pm 0.1$  by NaOH addition.  $\text{CaCl}_2$ , sodium phosphate and NaF were added to the final concentrations of 100 mM, 60 mM and 12 mM respectively. Following the addition of the calcium, phosphate and fluoride solutions the solution was microfiltered through a 0.1 or 0.2  $\mu\text{m}$  microfilter (ceramic or organic) to concentrate the solution five fold. The retentate was then diafiltered with one to five volumes of distilled water. The retentate after diafiltration was spraydried to produce a white powder that was 50% CPP and 40% ACFP and residue water.

The powdered CPP-ACFP was then reconstituted in distilled water to produce highly concentrated solutions. For example, a 10% w/v CPP-ACFP solution containing 640 mM Ca, 400 mM phosphate and 80 mM F (1,520 ppm F) at pH 9.0 has been prepared as well as a 20% CPP gel containing 1.28 M Ca, 800 mM phosphate and 160 mM F (3,040 ppm F) at pH 9.0. This solution and gel exhibit a significantly greater anticariogenicity relative to the fluoride alone and therefore are superior additives to toothpaste and mouthwash and for professional application to improve the efficacy of the current fluoride-containing dentifrices and professionally-applied products.

### Example 2 Structural Studies of CCP-ACP

#### *A. Structure and Interaction of CCP-ACP*

Casein phosphopeptides containing the Ser(P) cluster, ie. the core

sequence motif Ser(P)-Ser(P)-Ser(P)-Glu-Glu-, have a marked ability to stabilize calcium phosphate in solution. Solutions containing 0.1% w/v  $\alpha_{s1}$ (59-79) [1] at various pH, calcium and phosphate concentrations, but constant ionic strengths were used to characterise the peptide's interaction with calcium phosphate. The peptide was found to maximally bind 24 Ca and 16 Pi per molecule as shown in Table 1.

The ion activity products for the various calcium phosphate phases [hydroxyapatite (HA); octacalcium phosphate (OCP); tricalcium phosphate (TCP); amorphous calcium phosphate (ACP); and dicalcium phosphate dihydrate (DCPD)] were determined from the free calcium and phosphate concentrations at each pH using a computer program that calculates the ion activity coefficients through the use of the expanded Debye-Hückel equation and takes into account the ion pairs  $\text{CaHPO}_4^0$ ,  $\text{CaH}_2\text{PO}_4^+$ ,  $\text{CaPO}_4^-$  and  $\text{CaOH}^+$  the dissociation of  $\text{H}_3\text{PO}_4$  and  $\text{H}_2\text{O}$  and the ionic strength. The only ion activity product that significantly correlated with calcium phosphate bound to the peptide independently of pH was that corresponding to ACP [ $\text{Ca}_3(\text{PO}_4)_{1.87}(\text{HPO}_4)_{0.2}\text{xH}_2\text{O}$ ] indicating that this is the phase stabilised by  $\alpha_{s1}$ (59-79). The peptide  $\alpha_{s1}$ (59-79) binds to forming ACP clusters producing a metastable solution preventing ACP growth to the critical size required for nucleation and precipitation. The binding of  $\alpha_{s1}$ (59-79) to ACP results in the formation of colloidal complexes with the unit formula [ $\alpha_{s1}$ (59-79)(ACP) $_8$ ] $_n$  where n is equal to or greater than one. It is likely that the predominant form is n = 6 as  $\alpha_{s1}$ (59-79) cross-linked with glutaraldehyde in the presence of ACP runs as a hexamer on polyacrylamide gel electrophoresis. Interestingly, the synthetic octapeptide  $\alpha_{s1}$ (63-70) AcGlu-Ser(P)-Ile-Ser(P)-Ser(P)-Ser(P)-Glu-GluNHMe only binds 12 Ca and 8 Pi per molecule i.e. (ACP) $_4$  and the synthetic peptides corresponding to the N-terminus  $\alpha_{s1}$ (59-63), Gln-Met-Glu-Ala-Glu and the C-terminus  $\alpha_{s1}$ (71-78), Ile-Val-Pro-Asn-Ser(P)-Val-Glu-Gln of  $\alpha_{s1}$ (59-79) did not bind calcium phosphate as shown in Table 1. These results indicate that conformational specificity is essential for full ACP binding.

### B. NMR Studies

Protein flexibility in solution is the outstanding characteristic to emerge from spectroscopy studies on proteins containing the Ser(P) cluster sequence (-Ser(P)-Ser(P)-Ser(P)-) such as phosvitin from egg yolk and phosphophoryn from tooth dentine. Phosphorylation appears to destabilise secondary and tertiary structure rather than promote higher levels of ordering. However, flexible phosphorylated sequences adapt more regular conformations when bound to calcium phosphate. Optical rotatory dispersion (ORD), circular dichroism (CD), hydrodynamic and  $^{31}\text{P}$  nuclear magnetic resonance (NMR) measurements of the caseins all indicate that  $\alpha_{s1}$ -casein and  $\beta$ -casein have a rather open structure in solution with many amino acid side chains exposed to solvent and relatively flexible.  $^{31}\text{P}$ -NMR relaxation measurements indicate that Ser(P) residues are relatively mobile in  $\beta$ -casein.

We have demonstrated medium- and long-range nuclear Overhauser enhancements (nOes) in 2D  $^1\text{H}$  NMR spectra of  $\alpha_{s1}$ (59-79) [1] in the presence of  $\text{Ca}^{2+}$  indicating a conformational preference. Two structured regions were identified. Residues Val72 to Val76 are implicated in a  $\beta$ -turn conformation. Residues Glu61 to Ser(P)67, which extend over part of the Ser(P) cluster motif -Ser(P)-Ser(P)-Ser(P)-Glu-Glu- are involved in a loop-type structure. 2D NMR studies on  $\beta$ -casein(1-25) in the presence of calcium have shown a medium range nOe in the -Ser(P) $^{17}$ -Ser(P)-Ser(P)-Glu-Glu $^{21}$ - motif region between the C $\alpha$ H of Ser(P) $^{18}$  and NH of Glu $^{20}$ . Further medium range nOes include one between the C $\alpha$ H of Ser $^{22}$  and NH of Thr $^{24}$ . Evidence from the  $^1\text{H}$  NMR spectra of  $\alpha_2$ -casein(1-21) [4] have shown that several residues including those around the -Ser(P)-Ser(P)-Ser(P)-Glu-Glu- are perturbed. Furthermore, there are medium range nOes between NH of Ser(P) $^8$  and NH of Glu $^{10}$ . This is yet another example of a medium range nOe in the -Ser(P)-Ser(P)-Ser(P)-Glu-Glu-motif. Other examples of medium range nOes include that between the NH of Ile $^{14}$  and NH of Ser(P) $^{16}$ .



In summary the NMR data indicates that preferred conformations exist for these peptides in the presence of calcium ions. Molecular modeling of both  $\alpha_{s1}$ (59-79) and  $\beta$ (1-25) using the constraints derived from the NMR spectroscopy have indicated that the peptides adopt conformations that allow  
5 both glutamyl and phosphoseryl side chains of the cluster motif -Ser(*P*)-Ser(*P*)-Ser(*P*)-Glu-Glu to interact collectively with calcium ions of the ACP.

The relationship between CPP structure and interaction with amorphous calcium phosphate was investigated using a series of synthetic peptide homologues and analogues indicated in Table 1. These studies showed that the  
10 cluster sequence -Ser(*P*)-Ser(*P*)-Ser(*P*)-Glu-Glu- was mainly responsible for the interaction with ACP and that all three contiguous Ser(*P*) residues are required for maximal interaction with ACP.

TABLE 1.

**Calcium Phosphate Binding by CPP and Synthetic Homologues  
and Analogues**

	$V_{Ca}$ mol/mol	$V_{Pi}$ mol/mol	Ca/P
$\Sigma\Sigma\Sigma EE$	9	6	1.5
$S\Sigma\Sigma EE$	2	1	2.0
$E\Sigma\Sigma EE$	2	1	2.0
$D\Sigma\Sigma EE$	2	1	2.0
$\theta\theta\theta EE$	9	6	1.5
$S\theta\theta EE$	2	1	2.0
$A\Sigma AE$	0	0	
$IA\Sigma AEA$	0	0	
$EAIA\Sigma AEA$	0	0	
$A\Sigma A\Sigma AE$	0	0	
$A\Sigma A\Sigma A\Sigma AE$	2	1	1.5
$A\Sigma A\Sigma A\Sigma A\Sigma AE$	6	4	1.5
$\alpha_{s1}(59-79)$ QMEAE $\Sigma\Sigma\Sigma\Sigma\Sigma EE$ IVPN $\Sigma$ VEQK	24	16	1.5
$\alpha_{s1}(63-70)$ $\Sigma\Sigma\Sigma\Sigma\Sigma EE$	12	8	1.5
$\alpha_{s1}(66-70)$ $\Sigma\Sigma\Sigma EE$	9	6	1.5
$\alpha_{s1}(71-78)$ IVPN $\Sigma$ VEQK	0	0	
$\alpha_{s1}(59-63)$ QMEAE	0	0	
$\beta(1-25)$ RELEELNVPGEIV $\Sigma$ E $\Sigma\Sigma\Sigma\Sigma$ EESITR	24	16	1.5
$\beta(14-21)$ $\Sigma\Sigma\Sigma\Sigma\Sigma EE$	12	8	1.5

$\Sigma$ =Ser(P),  $\theta$ =Thr(P), E = Glu, D = Asp, S = Ser, A = Ala, I = Ile, Q = Gln,  
M = Met, V = Val, P = Pro, K = Lys, L = Leu, T = Thr, G = Gly and R = Arg.

**Example 3 Structural Studies Using Hydroxyapatite (HA)**

Similarly, we investigated the adsorption of the CPP and synthetic homologues and analogues onto HA (Table 2). These data also confirm that the Ser(*P*) cluster sequence is the major determinant for high affinity binding and  
5 that all three contiguous Ser(*P*) residues are essential as loss of any one, even when substituted with a Glu or Asp, resulted in a considerably lower affinity constant K as shown in Table 2.

TABLE 2.  
 CPP and Synthetic Peptide binding to HA at 37°C

	K ml/ μmol	N μmol/ m <sup>2</sup>	Molecular Area nm <sup>2</sup>
α <sub>s1</sub> (59-79) QMEAEΣΙΣΣΣEEIVPNΣVEQK	415	0.35	4.75
α <sub>s1</sub> (63-70) EΣΙΣΣΣEE	10,370	0.47	3.56
α <sub>s1</sub> (66-70) ΣΣΣEE	12,845	0.52	3.27
α <sub>s1</sub> (71-78) IVPNΣVEQK	-	-	-
α <sub>s1</sub> (59-63) QMEAE	-	-	-
ΣΣΣEE	12,845	0.52	3.27
EΣΣEE	1,513	0.96	1.74
DΣΣEE	6,579	0.81	2.04
θθθEE	12,234	0.51	3.27
TθθEE	1,013	0.55	3.03
θTθEE	837	0.44	3.77
θθTEE	1,799	0.46	3.61

Σ = Ser(P), θ = Thr(P)

Interestingly, repeating these HA adsorption experiments with salivary coated HA (sHA) revealed that the Ser(P) cluster motif was still the major determinant for adsorption although the affinities of the peptides for the sHA was slightly reduced by the presence of the salivary proteins. These results suggest that the predominant interaction of the CPP with pellicle and plaque is likely to be electrostatic and mediated by the Ser(P) cluster motif of the CPP.

We have also studied the docking of the peptide Ser(P)-Ser(P)-Ser(P)-Glu-Glu- onto three crystallographic planes of HA, {100}, {010} and {001} using computer simulation techniques and the unit cell coordinates of synthetic HA. These simulation studies revealed that the peptide -Ser(P)-Ser(P)-Ser(P)-Glu-Glu- is more likely to bind to the {100} surface, followed by the {010} surface. The Ser(P) cluster motif can therefore bind to both {100} and {010} surfaces thus allowing deposition of calcium, phosphate and hydroxyl ions on the {001} surface enabling growth of the HA crystal along the *c*-axis only. These results therefore can now explain the *c*-axis growth of HA crystals in enamel and dentine. Detailed examination of the computer simulation data shows that the -Ser(P)-Ser(P)-Ser(P)-Glu-Glu- conformer with the greatest relative binding energy is positioned on the HA surface such that the carboxyl groups of the glutamyl residues and the phosphoryl groups of the phosphoserine residues are in proximity to the HA surface with maximal contact between these groups and surface calcium atoms.

#### **Example 4 Anticariogenic Activity of CPP-ACP in Human *in situ* Studies**

The ability of the 1.0% w/v CPP-ACP pH 7.0 solution to prevent enamel demineralisation was studied in a human *in situ* caries model. The model consists of a removable appliance containing a left and right pair of enamel slabs placed to produce a plaque retention site. The inter-enamel plaque that developed (3-5 mg) was bacteriologically similar to normal supragingival plaque. On frequent exposure to sucrose solutions over a three week period, the increase in levels of mutans streptococci and lactobacilli and in sub-surface

enamel demineralisation resulted in the formation of incipient "caries-like" lesions.

Two exposures of the CPP-ACP solution per day to the right pair of enamel slabs for 12 subjects produced a  $51\% \pm 19\%$  reduction in enamel mineral loss relative to the left-side, control enamel. The plaque exposed to the CPP-ACP solution contained  $78 \pm 22 \mu\text{mol/g}$  calcium,  $52 \pm 25 \mu\text{mol/g}$   $\text{P}_i$  and  $2.4 \pm 0.7 \text{ mg/g}$  CPP compared with  $32 \pm 12 \mu\text{mol/g}$  calcium and  $20 \pm 11 \mu\text{mol/g}$   $\text{P}_i$  in the control plaque. The level of the CPP was determined by competitive ELISA using an antibody that recognises both  $\alpha_1(59-79)$  and  $\beta(1-25)$ . Electron micrographs of immunocytochemically stained sections of the plaque revealed localisation of the peptide predominantly on the surface of microorganisms but also in the extracellular matrix.

Although these results indicate that CPP are incorporated into developing dental plaque, the actual level determined by ELISA would not be a true representation of that incorporated due to the breakdown of the CPP in plaque through the action of phosphatase and peptidase activities. The incorporation of the CPP-ACP in the plaque resulted in a 2.4 fold increase in the plaque calcium and a 2.6 fold increase in plaque  $\text{P}_i$  with a  $\text{Ca}/\text{P}_i$  ratio consistent with ACP.

#### **Example 5 Anticariogenic Potential of the CPP-ACP in a Mouthwash Study**

A clinical trial of a mouthwash used thrice daily containing 3.0% CPP-ACP pH 9.0 showed that the calcium content of supragingival plaque (lower anterior teeth excluded) increased from  $169 \pm 103 \mu\text{mol/g}$  dry weight to  $610 \pm 234 \mu\text{mol/g}$  after use of the mouthwash for a three day period, and inorganic phosphate increased from  $242 \pm 60 \mu\text{mol/g}$  dry weight to  $551 \pm 164 \mu\text{mol/g}$ . These post-mouthwash levels of calcium and inorganic phosphate are the highest ever reported for non-mineralised supragingival plaque.

Without wishing to be bound by any proposed mechanism for the observed advantages, it is believed that the mechanism of anticariogenicity for

the CPP-ACP is the incorporation of amorphous calcium phosphate in plaque, thereby depressing enamel demineralisation and enhancing remineralisation. In plaque, CPP-ACP would act as a reservoir of calcium and phosphate, buffering the free calcium and phosphate ion activities thereby helping to maintain a state of supersaturation with respect to tooth enamel. The binding of ACP to CPP is pH dependent with very little bound below pH 7.0.

### **Example 6 Remineralisation of Enamel Lesions by CPP-ACP**

#### ***A. In vitro studies***

An *in vitro* enamel remineralisation system was used to study remineralisation of artificial lesions in human third molars by CPP-ACP solutions. Using this system, a 1.0% CPP-ACP solution replaced  $56 \pm 21\%$  of mineral lost. A 0.1% CPP-ACP solution replaced  $34 \pm 18\%$  of mineral lost. A further number of solutions containing various amounts of CPP (0.1-1.0%), calcium (6-60 mM) and phosphate (3.6-36 mM) at different pH values (7.0-9.0) were prepared. The associations between the activities of the various calcium phosphate species in solution and the rate of enamel lesion remineralisation for this series of solutions were then determined.

The activity of the neutral ion species  $\text{CaHPO}_4^0$  in the various remineralising solutions was found to be highly correlated with the rate of lesion remineralisation. The diffusion coefficient for the remineralisation process was estimated at  $3 \times 10^{-10} \text{ m}^2\text{s}^{-1}$  which is consistent with the coefficients of diffusion for neutral molecules through a charged matrix. The rate of enamel remineralisation obtained with the 1.0% CPP-ACP solution was  $3.3 \times 10^{-2} \text{ mol HA/m}^2/10 \text{ days}$  which is the highest remineralisation rate ever obtained. Calcium phosphate ions, in particular the neutral ion pair  $\text{CaHPO}_4^0$ , after diffusion into the enamel lesion, will dissociate and thereby increase the degree of saturation with respect to HA. The formation of HA in the lesion will lead to the generation of  $\text{H}_3\text{PO}_4$ , which being neutral itself, will diffuse out of the lesion down a concentration gradient.

The results indicate that the CPP-bound ACP,  $\text{CPP}[\text{Ca}_3(\text{PO}_4)_{1.87}(\text{HPO}_4)_{0.2}\text{xH}_2\text{O}]_8$  acts as a reservoir of the neutral ion species,  $\text{CaHPO}_4^0$  that is formed in the presence of acid. The acid can be generated by dental plaque bacteria; under these conditions, the CPP-bound ACP would buffer plaque pH and produce calcium and phosphate ions, in particular  $\text{CaHPO}_4^0$ . The increase in plaque  $\text{CaHPO}_4^0$  would offset any fall in pH thereby preventing enamel demineralisation. Acid is also generated in plaque as  $\text{H}_3\text{PO}_4$  by the formation of HA in the enamel lesion during remineralisation. This therefore explains why the CPP-ACP solutions are such efficient remineralising solutions as they would consume the  $\text{H}_3\text{PO}_4$  produced during enamel lesion remineralisation generating more  $\text{CaHPO}_4^0$  thus maintaining its concentration gradient into the lesion. These results are therefore consistent with the proposed anticariogenic mechanism of the CPP being the inhibition of enamel demineralisation and enhancement of remineralisation through the localisation of ACP at the tooth surface.

*B. Human in situ remineralisation studies*

The ability of CPP-ACP added to sugar-free (sorbitol) chewing gum to remineralise enamel sub-surface lesions was investigated in a randomized, cross-over, double-blind study. Ten subjects wore removable palatal appliances with six, human-enamel, half-slabs inset containing sub-surface demineralised lesions. The other half of each enamel slab was stored in a humidified container and was used as the control demineralised lesion. There were four treatment groups in the study, sugar-free gum containing 3.0% w/w CPP-ACP, sugar-free gum containing 1.0% w/w CPP-ACP, sugar-free gum with no CPP-ACP and a no-gum-chewing control. The gums were chewed for 20 min periods, four times a day. The appliances were worn for this 20 min period and a further 20 min period after gum chewing. Each treatment was for 14 days duration and each of the ten subjects carried out each treatment with a one week rest between the treatments. At the completion of each treatment the enamel slabs were removed, paired with their respective demineralised control, embedded,



sectioned and subjected to microradiography and computer-assisted densitometric image analysis to determine the level of remineralisation. The sugar-free gum treatment resulted in  $9.82 \pm 1.81\%$  remineralisation relative to the no-gum-chewing control whereas the gum containing 1.0% CPP-ACP produced  $17.06 \pm 2.48\%$  remineralisation and the 3.0% CPP-ACP gum produced  $22.70 \pm 3.40\%$  remineralisation with all values being significantly different. These results showed that addition of 1.0% and 3.0% CPP-ACP to sugar-free gum produced a 74% and 131% increase respectively in sub-surface enamel remineralisation.

10 **Example 7 CPP-ACFP Mouthwash Study**

A mouthwash study was conducted to determine the ability of a 3.0% CPP-ACFP mouthwash used thrice daily to increase supragingival plaque calcium, inorganic phosphate and fluoride ions. The 3.0% CPP-ACFP solution used as a mouthwash for four days contained 192 mM bound calcium ions, 120 mM bound phosphate ions and 24mM (456 ppm) bound F ions stabilised by CPP. The use of the mouthwash resulted in a 1.9 fold increase in plaque calcium, a 1.5 fold increase in plaque phosphate and a dramatic 18 fold increase in plaque fluoride ion as shown in Table 3.

**TABLE 3.**  
**Effect of CPP-ACFP on Plaque, Ca, P<sub>i</sub> and F Levels**

	Ca μmol/g	Pi μmol/g	F μmol/g
Control	177 ± 53	306 ± 82	1.1 ± 0.9
3% CPP - ACFP	336 ± 107	471 ± 113	19.9 ± 14.1
1000 ppm F	158 ± 54	287 ± 29	1.9 ± 1.0
3% CaCPP	193 ± 56	343 ± 102	1.5 ± 0.8

Although these marked increases in plaque calcium, phosphate and fluoride were found, dental calculus was not observed in any of the subjects, suggesting that the plaque calcium fluoride phosphate remained stabilised as the amorphous phase by the CPP and did not transform into a crystalline phase.

- 5 These increases in the supragingival plaque levels of Ca, phosphate and fluoride ions produced by CPP-ACP are markedly greater than those obtained in a similar study using CaCPP and 1000 ppm F (MFP and NaF) toothpastes twice daily for a similar time period as indicated in Table 3. These results show a marked synergistic effect between fluoride ions and the CPP-ACP. This is particularly advantageous in view of the fact that the level of fluoride in oral compositions such as toothpaste can then be reduced, resulting in cost savings and lowered risk of fluorosis for individuals living in high-fluoride areas.

#### **Example 8 Interaction of CPP-ACP with Fluoride**

- 15 An synergistic anticariogenic effect of the 1.0% CPP-ACP together with 500 ppm F<sup>-</sup> was observed in a rat caries model. Analysis of the solution containing 1.0% CPP, 60 mM CaCl<sub>2</sub>, 36 mM sodium phosphate and 500 ppm F (26.3 mM NaF) pH 7.0 after ultrafiltration revealed that nearly half of the fluoride ion had incorporated into the ACP phase stabilised by the CPP to produce an amorphous calcium fluoride phosphate phase of composition
- 20  $\text{Ca}_3(\text{PO}_4)_2\text{F}_x\text{H}_2\text{O}$ , with 24 Ca, 15 PO<sub>4</sub> and 3F molecules per CPP molecule.

- Without wishing to be limited by any proposed mechanism for the observed beneficial effect, we consider that the anticariogenic mechanism of the CPP-ACP is the localisation of ACP at the tooth surface such that in the presence of acid, the ACP dissociates to release Ca and phosphate ions
- 25 increasing the degree of saturation with respect to HA preventing enamel demineralisation and promoting remineralisation. The anticariogenic mechanism of fluoride is the localisation of the fluoride ion at the tooth surface, particularly in plaque in the presence of Ca and phosphate ions. This localisation increases the degree of saturation with respect to fluorapatite (FA) thus promoting

remineralisation of enamel with FA. It is clear that for the formation of FA  $[\text{Ca}_{10}(\text{PO}_4)_6\text{F}_2]$ , calcium and phosphate ions must be co-localised in plaque at the tooth surface with the fluoride ion. The synergistic anticariogenic effect of CPP-ACP and F is therefore attributable to the localisation of ACFP at the tooth surface by the CPP which in effect would co-localise Ca, Pi and F.

This was demonstrated in the mouthwash study described in Example 7.

Metastable solutions of the CPP at pH 7.0 have been prepared containing amorphous calcium fluoride phosphate at remarkably high concentrations. For example, a 10% w/v CPP -ACFP solution containing 640 mM Ca, 400 mM phosphate and 80 mM F (1,520 ppm F<sup>-</sup>) at pH 7.0 has been prepared as well as a 20% CPP gel containing 1.28 M Ca, 800 mM phosphate and 160 mM F (3,040 ppm F<sup>-</sup>) at pH 7.0. This solution and gel exhibit a significantly greater anticariogenicity relative to the fluoride alone, and therefore are superior additives to toothpastes and mouthwash and for professional application to improve the efficacy of the current fluoride-containing dentifrices and professionally-applied products.

Specific examples of formulations containing the complexes of the invention are provided below.

Example 9 Toothpaste formulations containing CPP-ACFP

Formulation 1

<u>Ingredient</u>	<u>% w/w</u>
Dicalcium phosphate dihydrate	50.0
Glycerol	20.0
Sodium carboxymethyl cellulose	1.0
25 Sodium lauryl sulphate	1.5
Sodium lauroyl sarconisate	0.5
Flavour	1.0
Sodium saccharin	0.1
Chlorhexidine gluconate	0.01
30 Dextranase	0.01

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CPP-ACFP	1.00
Water	balance

*Formulation 2*

	<u>Ingredient</u>	<u>% w/w</u>
5	Dicalcium phosphate dihydrate	50.0
	Sorbitol	10.0
	Glycerol	10.0
	Sodium carboxymethyl cellulose	1.0
	Sodium lauryl sulphate	1.5
10	Sodium lauroyl sarconisate	0.5
	Flavour	1.0
	Sodium saccharin	0.1
	Sodium monofluorophosphate	0.3
	Chlorhexidine gluconate	0.01
15	Dextranase	0.01
	CPP-ACFP	2.0
	Water	balance

*Formulation 3*

	<u>Ingredient</u>	<u>% w/w</u>
20	Dicalcium phosphate dihydrate	50.0
	Sorbitol	10.0
	Glycerol	10.0
	Sodium carboxymethyl cellulose	1.0
	Lauroyl diethanolamide	1.0
25	Sucrose monolaurate	2.0
	Flavour	1.0
	Sodium saccharin	0.1
	Sodium monofluorophosphate	0.3

29

Chlorhexidine gluconate	0.01
Dextranase	0.01
CPP-ACFP	5.0
Water	balance

5 *Formulation 4*

	<u>Ingredient</u>	<u>% w/w</u>
	Sorbitol	22.0
	Irish moss	1.0
	Sodium Hydroxide (50%)	1.0
10	Gantrez	19.0
	Water (deionised)	2.69
	Sodium Monofluorophosphate	0.76
	Sodium saccharine	0.3
	Pyrophosphate	2.0
15	Hydrated alumina	48.0
	Flavour oil	0.95
	CPP-ACFP	1.0
	sodium lauryl sulphate	2.00

*Formulation 5*

	<u>Ingredient</u>	<u>% w/w</u>
20	Sodium polyacrylate	50.0
	Sorbitol	10.0
	Glycerol	20.0
	Flavour	1.0
25	Sodium saccharin	0.1
	Sodium monofluorophosphate	0.3
	Chlorhexidine gluconate	0.01
	Ethanol	3.0

30

CPP-ACFP	2.0
Linolic acid	0.05
Water	balance

Example 10 Mouthwash formulations5 *Formulation 1*

	<u>Ingredient</u>	<u>% w/w</u>
	Ethanol	20.0
	Flavour	1.0
	Sodium saccharin	0.1
10	Sodium monofluorophosphate	0.3
	Chlorhexidine gluconate	0.01
	Lauroyl diethanolamide	0.3
	CPP-ACFP	2.0
	Water	balance

15 *Formulation 2*

	<u>Ingredient</u>	<u>% w/w</u>
	Gantrez S-97	2.5
	Glycerine	10.0
	Flavour oil	0.4
20	Sodium monofluorophosphate	0.05
	Chlorhexidine gluconate	0.01
	Lauroyl diethanolamide	0.2
	CPP-ACFP	2.0
	Water	Balance

Example 11 Lozenge formulation

	<u>Ingredient</u>	<u>% w/w</u>
	Sugar	75-80
	Corn syrup	1-20
5	Flavour oil	1-2
	NaF	0.01-0.05
	CPP-ACFP	3.0
	Mg stearate	1-5
	Water	balance

10 Example 12 Gingival Massage Cream formulation

	<u>Ingredient</u>	<u>% w/w</u>
	White petrolatum	8.0
	Propylene glycol	4.0
	Stearyl alcohol	8.0
15	Polyethylene Glycol 4000	25.0
	Polyethylene Glycol 400	37.0
	Sucrose monostearate	0.5
	Chlorohexidine gluconate	0.1
	CPP-ACFP	3.0
20	Water	balance

Example 13 Chewing Gum formulation

	<u>Ingredient</u>	<u>% w/w</u>
	Gum base	30.0
	Calcium carbonate	2.0
25	Crystalline sorbitol	53.0
	Glycerine	0.5
	Flavour oil	0.1
	CPP-ACFP	2.0
	Water	balance



Example 11 Dietary Supplement

5 CPP-ACP was added at 1.0% w/w of the diet of rachitic chickens to determine the ability of the CPP-ACP to provide bioavailable calcium for bone accretion. CPP-ACP at 1.0% w/w in the diet produced a 34% reduction in the incidence of growth plate abnormalities, a 17% increase in tibial ash and a 22% reduction in the cartilaginous growth plate in the animals which was significantly greater than the CPP alone (Table 4) indicating that the CPP-ACP is superior to the CPP in providing bioavailable dietary calcium and in facilitating bone accretion.

TABLE 4.

Effect of 1.0% CPP-ACP addition to the diet of rachitic chickens  
on incidence of growth plate abnormalities, tibial ash and  
cartilaginous growth plate width

	% Growth Abnormalities %	% Tibial Ash %	Growth Plate Width (mm)
Control	53 $\pm$ 5	30 $\pm$ 2	5.4 $\pm$ 0.2
1.0% CPP	47 $\pm$ 9	30 $\pm$ 2	5.3 $\pm$ 0.2
1.0% CPP-ACP	35 $\pm$ 3	35 $\pm$ 1	4.2 $\pm$ 0.2

It should be understood that while the invention has been described in detail for the purposes of clarity and understanding, the examples were for illustrative purposes only. Other modifications of the embodiments of the present invention will be apparent to those skilled in the art of molecular biology, dental diagnostics, and related disciplines and are within the scope of the invention as described.

## CLAIMS

1. A stable calcium phosphate complex including phosphopeptide stabilized amorphous calcium phosphate or a derivative thereof wherein said phosphopeptide includes the amino acid sequence Ser(P)-Ser(P)-Ser(P)-Glu-Glu.
- 5 2. A complex according to claim 1 wherein said amorphous calcium phosphate is of the formula  $[\text{Ca}_3(\text{PO}_4)_{1.87}(\text{HPO}_4)_{0.2}x\text{H}_2\text{O}]$  wherein  $x \geq 1$ .
3. A complex according to claim 1 wherein said amorphous calcium phosphate derivative is calcium fluoride phosphate of the formula  $[\text{Ca}_8(\text{PO}_4)_5 \text{ F } x \text{ H}_2\text{O}]$  wherein  $x \geq 1$ .
- 10 4. A complex according to any one of claims 1 to 3 wherein said phosphopeptide includes an amino acid sequence selected from any one of:
  - [1] Gln<sup>59</sup>-Met-Glu-Ala-Glu-Ser(P)-Ile-Ser(P)-Ser(P)-Ser(P)-Glu-Glu-Ile-Val-Pro-Asn-Ser(P)-Val-Glu-Gln-Lys<sup>79</sup>.  $\alpha_{51}(59-79)$
  - [2] Arg<sup>1</sup>-Glu-Leu-Glu-Glu-Leu-Asn-Val-Pro-Gly-Glu-Ile-Val-Glu-Ser(P)-Leu-Ser(P)-Ser(P)-Ser(P)-Glu-Glu-Ser-Ile-Thr-Arg<sup>25</sup>.  $\beta(1-25)$
  - 15 [3] Asn<sup>46</sup>-Ala-Asn-Glu-Glu-Glu-Tyr-Ser-Ile-Gly-Ser(P)-Ser(P)-Ser(P)-Glu-Glu-Ser(P)-Ala-Glu-Val-Ala-Thr-Glu-Glu-Val-Lys<sup>70</sup>.  $\alpha_{22}(46-70)$
  - [4] Lys<sup>1</sup>-Asn-Thr-Met-Glu-His-Val-Ser(P)-Ser(P)-Ser(P)-Glu-Glu-Ser-Ile-Ile-Ser(P)-Gln-Glu-Thr-Tyr-Lys<sup>21</sup>.  $\alpha_{22}(1-21)$ .
- 20 5. A complex according to any one of claims 1 to 3 wherein said phosphopeptide includes the amino acid sequence:
 

Gln<sup>59</sup>-Met-Glu-Ala-Glu-Ser(P)-Ile-Ser(P)-Ser(P)-Ser(P)-Glu-Glu-Ile-Val-Pro-Asn-Ser(P)-Val-Glu-Gln-Lys<sup>79</sup>.  $\alpha_{51}(59-79)$ .
- 25 6. A complex according to any one of claims 1 to 5 having the formula  $[(\text{PP})(\text{CP})_8]_n$  where n is equal to or greater than 1.
7. A complex according to any one of claims 1 to 6 having the formula  $[(\text{PP})(\text{CP})_8]_6$ .
8. A delivery vehicle adapted to co-localise ions including calcium, fluoride and phosphate ions at a target site, said delivery vehicle including as
 

30 an active ingredient, the complex of any one of claims 1 to 7.

9. A delivery vehicle of claim 8 being any one of toothpaste, toothpowder, liquid dentifrice, mouthwash, troches, chewing gum, dental paste, gingival massage cream, gargle tablets, dairy products and other foodstuffs.

10. A pharmaceutical composition including a complex according to  
5 any one of claims 1 to 7 in combination with a pharmaceutical carrier.

11. An anticariogenic composition including the composition of claim 10 or the delivery vehicle according to claim 8 or 9 adapted for anticariogenic use.

12. A therapeutic composition including the composition of claim 10  
10 adapted for therapeutic use.

13. A dietary supplement including a complex according to any one of claims 1 to 7.

14. A dietary supplement according to claim 13 wherein said complex comprises 0.1 to 100% w/w of said supplement.

15. 15. A dietary supplement according to claim 13 wherein said complex comprises 2% w/w of said supplement.

16. A method of treating or preventing dental caries or tooth decay including the step of administering a complex according to any one of claims 1 to 7 to the teeth or gums of a subject in need of such treatment.

20 17. A method according to claim 16 including the topical administration of said complex.

18. A method of treating one or more conditions related to calcium loss from the body including bones, calcium deficiency, calcium malabsorption or the like including the administration of a complex according to any one of  
25 claims 1 to 7 to a subject in need of such treatment.

19. The method of claim 18 wherein said condition includes osteoporosis and osteomalacia.

20. A method of producing a stable complex of calcium phosphate according to any one of claims 1 to 7 including the steps of:

30 (i) obtaining a solution of phosphopeptide having a pH of about 9.0;

- (ii) admixing (i) with solutions comprising calcium, and inorganic phosphate and optionally fluoride at a pH of about 9.0;
  - (iii) filtering the mixture resulting from step (ii), and
  - (iv) drying to obtain the said complex.
- 5        21. A complex according to any one of claims 1 to 7 as hereinbefore described.
22. A delivery vehicle according to claim 8 or 9 as hereinbefore described.
23. A pharmaceutical composition according to claim 10 as  
10 hereinbefore described.
24. An anticariogenic composition according to claim 11 as hereinbefore described.
25. A therapeutic composition according to claim 12 as hereinbefore described.
- 15        26. A dietary supplement according to any one of claims 13 to 15 as hereinbefore described.
27. A method of treatment according to any one of claims 16 to 19 substantially as hereinbefore described.
28. A method according to claim 20 substantially as hereinbefore  
20 described.

## INTERNATIONAL SEARCH REPORT

International Application No.  
PCT/AU 98/00160

C (Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT		
Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	Biochimica et Biophysica Acta, volume 1247, 1995 N Huq et al, "A <sup>1</sup> H-NMR Study of the Casein Phosphopeptide $\alpha_{11}$ -casein (59-79)," pages 201-208 entire document	1-20
X	Journal of Dental Research, volume 74, no: 6, 1995 E C Reynolds et al, "Anticariogenicity of Calcium Phosphate Complexes of Tryptic Casein Phosphopeptides in the Rat," pages 1272-1279 entire document (see abstract lines 1-3)	1-20
X	Biochemical Journal, volume 314, no: 3, 1996 C Holt et al, "Ability of a $\beta$ -casein Phosphopeptide to Modulate the Precipitation of Calcium Phosphate by Forming Amorphous Dicalcium Phosphate Nanoclusters," pages 1035-1039 entire document, especially page 1036 column 2	1-20
X	Journal of Chromatography, volume 646, 1993 E C Reynolds et al, "The Analysis of Multiple Phosphoserine-containing Casein Peptides Using Capillary Zone Electrophoresis," pages 391-396 abstract and page 391 first paragraph	1-20
X	Enzyme and Microbial Technology, volume 19, 1996 N J Adamson and E C Reynolds, "Characterization of Casein Phosphopeptides Prepared using Alcalase: Determination of Enzyme Specificity," pages 202-207 abstract and page 202 second column	1-20
P,X	Journal of Dental Research, volume 76, no: 9, 1997 E C Reynolds, "Remineralization of Enamel Subsurface Lesions by Casein Phosphopeptide-Stabilized Calcium Phosphate Solutions," pages 1587-1595 entire document, especially abstract and page 1588 left column, third paragraph	1-20
A	WO 82/03008 (THE UNIVERSITY OF MELBOURNE et al), 16 September 1982 entire document	1-20
A	WO 87/07615 (THE UNIVERSITY OF MELBOURNE et al), 17 December 1987 entire document	1-20
A	Archives of Oral Biology, volume 39, no: 8 1994 K Wilkiel et al, "Hydroxyapatite Mineralization and Demineralization in the Presence of Synthetic Phosphorylated Pentapeptides," pages 715-721 entire document	1-20
A	WO 96/29340 (HAN SANG KEE), 26 September 1996 entire document	1-20

# INTERNATIONAL SEARCH REPORT

International Application No.  
PCT/AU 98/00160

<b>A. CLASSIFICATION OF SUBJECT MATTER</b>		
Int Cl <sup>6</sup> : C07K 14/47, C07K 101.02; C01B 25/32; A61K 7/16, 33/42, 38/08, 38/16, 38/17		
According to International Patent Classification (IPC) or to both national classification and IPC		
<b>B. FIELDS SEARCHED</b>		
Minimum documentation searched (classification system followed by classification symbols) IPC as above		
Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched		
Electronic data base consulted during the international search (name of data base and, where practicable, search terms used) STN: CASA (i) sssee/sqsp and phosph?; and calcium (ii) casein and calcium (w) phosph?; and complex; and dent?		
<b>C. DOCUMENTS CONSIDERED TO BE RELEVANT</b>		
Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	WO 94/00146 (THE UNIVERSITY OF MELBOURNE et al) 6 January 1994 entire document, especially pages 3 to 5	1-20
X	Archives of Oral Biology, volume 36, no: 1, 1991 S HIDAKA et al, "A New Method for the Study of the Formation and Transformation of Calcium Phosphate Precipitates: Effects of Several Chemical Agents and Chinese Folk Medicines," pages 49-54 entire document, especially page 50 column 2 last paragraph	1-20
X	Proceedings of the Nutrition Society of Australia, 1995 E C Reynolds, "Dairy Products and Dental Health," pages 95-102 see especially page 98 second paragraph, page 100 line 17	1-20
<input checked="" type="checkbox"/> Further documents are listed in the continuation of Box C <input checked="" type="checkbox"/> See patent family annex		
* Special categories of cited documents: "A" document defining the general state of the art which is not considered to be of particular relevance "E" earlier document but published on or after the international filing date "L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified) "O" document referring to an oral disclosure, use, exhibition or other means "P" document published prior to the international filing date but later than the priority date claimed "T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention "X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone "Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art "&" document member of the same patent family		
Date of the actual completion of the international search 16 April 1998		Date of mailing of the international search report <b>22 APR 1998</b>
Name and mailing address of the ISA/AU AUSTRALIAN PATENT OFFICE PO BOX 200 WODEN ACT 2606 AUSTRALIA Facsimile No.: (02) 6285 3929		Authorized officer  O.L. CHAI <i>James Rider</i> Telephone No.: (02) 6283 2482



# INTERNATIONAL SEARCH REPORT

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## Box I Observations where certain claims were found unsearchable (Continuation of item 1 of first sheet)

This International Search Report has not been established in respect of certain claims under Article 17(2)(a) for the following reasons:

1. ☐ Claims Nos.:  
because they relate to subject matter not required to be searched by this Authority, namely:
2. ☐ Claims Nos.:  
because they relate to parts of the international application that do not comply with the prescribed requirements to such an extent that no meaningful international search can be carried out, specifically:
3. ☒ Claims Nos.: 21 to 28  
because they are dependent claims and are not drafted in accordance with the second and third sentences of Rule 6.4(a)

## Box II Observations where unity of invention is lacking (Continuation of item 2 of first sheet)

This International Searching Authority found multiple inventions in this international application, as follows:

1. ☐ As all required additional search fees were timely paid by the applicant, this international search report covers all searchable claims
2. ☐ As all searchable claims could be searched without effort justifying an additional fee, this Authority did not invite payment of any additional fee.
3. ☐ As only some of the required additional search fees were timely paid by the applicant, this international search report covers only those claims for which fees were paid, specifically claims Nos.:
4. ☐ No required additional search fees were timely paid by the applicant. Consequently, this international search report is restricted to the invention first mentioned in the claims; it is covered by claims Nos.:

### Remark on Protest

- ☐ The additional search fees were accompanied by the applicant's protest.  
☐ No protest accompanied the payment of additional search fees.



PCT

## INTERNATIONAL SEARCH REPORT

### Information on patent family members

International Application No.  
PCT/AU 98/00160

This Annex lists the known "A" publication level patent family members relating to the patent documents cited in the above-mentioned international search report. The Australian Patent Office is in no way liable for these particulars which are merely given for the purpose of information.

Patent Document Cited in Search Report				Patent Family Member			
WO	94/00146	AU	43006/93	EP	651652	NZ	252959
WO	82/03008	AU	81489/82	CA	1200506	CA	1212627
		EP	73210	NZ	211745	US	5130123
		NZ	199891				
WO	87/07615	AU	75483/87	BR	8707352	CA	1315480
		DK	703/88	EP	268663	FI	880629
		FI	931988	HU	47309		
WO	96/29340	AU	51246/96	CA	2216250	EP	815130
RECEIVED AUG 10 1999 TC 1700 MAIL ROOM							
END OF ANNEX							